BOARD OF INTERMEDIATE & SECONDARY EDUCATION, HYDERABAD

Excellence – Equity – Empathy

Time: 30 Minutes CHEMISTRY-I MODEL PAPER CLASS: IX Marks: 30

SECTION "A"

MULTIPLE CHOICE QUESTIONS (MCQ's)

Q.No:1 **Choose the correct answers for each from the given options.**

1. How many moles of sodium are present in 46g of substance?
   (a) 1.8 moles (b) 2.3 moles (c) 2 moles (d) 1mole

2. Those Acids which contains one acidic hydrogen are called:
   (a) Mono-protic (b) Di-protic (c) Tri-protic (d) polyprotic

3. Mg(OH)Cl is known as:
   (a) Acidic salt (b) Basic salt (c) Neutral salt (d) Normal salt

4. The formula of potassium chlorate:
   (a) KCO_2 (b) $KClO_3$ (c) KCl (d) $NaHCO_3$

5. The number of moles of solute dissolved in one dm^3 is called:
   (a) Molality (b) Avogadro's number (c) Morality (d) Mole

6. The shared pair of electrons which links the atoms in a molecule is known as:
   (a) Electro-valent bond (b) Covalent bond
   (c) Co-ordinate covalent bond (d) chemical bond

7. The only liquid metals:
   (a) Molybdenum (b) Gold (c) Mercury (d) Bromine

8. To which family Gabelongs:
   (a) Boron (b) Carbon (c) Nitrogen (d) Fluorine

9. Lothar Mayer's curve included about elements:
   (a) Thirty (b) Forty (c) Fifty six (d) Sixty two

10. The Most reactive metals:
    (a) Na (b) Fe (c) Cu (d) Ca

11. The PH of 0.01 M of HCl is:
    (a) 2 (b) 1 (c) 3 (d) None of these

12. Faraday's First law of electrolysis _____ is directly proportional to the amount of substance that deposited at any electrode:
    (a) Current (b) Weight (c) both of these (d) None of these

13. Which type of bond is present in Chlorine molecule:
    (a) polar covalent (b) non-polar covalent (c) Electrovalent (d) Dative bond.

14. Robert Brown used which powder for explaining the movement of particles:

(a) Sulphur (b) Phosphorus (c) Silicon (d) Magnesium

15. The one sided sharing of electrons is known as:

(a) Dative bond (b) Co-ionic (c) Co-ordinate covalent (d) All of these

16. Those substance which conduct electricity are known as:

(a) Electrolytes (b) non-electrolytes (c) Both of these (d) None of these

17. H_3O^+ is called:

(a) Hydronium ion (b) H^+ is Acceptor (c) O is Donor (d) All of these.

18. 60 g of NaOH dissolved in one litre is known as:

(a) 1.5 M solution (b) 2 M solution (c) 2.5 M solution (d) None of these.

19. Bitter Taste is representation of:

(a) Acid (b) Base (c) Salt (d) All of these

20. Nucleus is made up dense particle explained by:

(a) Rutherford (b) Bohr (c) Dalton (d) Al-Razi

21. Acids react's with base to neutralize and form:

(a) salt (b) Water (c) salt and water (d) None of these

22. Gaseous Directly converted into solid is:

(a) Desublimation (b) De-position (c) Sublimation (d) Both a & b

23. In combustion reaction we burn methane it forms:

(a) carbon dioxide (b) water (c) Hydrogen (d) Cabondioxide and water

24. Which type of bond is indicated by arrow:

(a) Dative (b) Ionic Bond (c) Electro-valent (d) Covalent

25. How many electron Ag losses in $AgNO^3$:

(a) one (b) two (c) three (d) four

26. W= ZAt is mathematical representation for:

(a) Faraday's first law (b) Faraday's 2nd law (c) Both of these (d) none of these

27. Water from different sources always in same proportion in known as:

(a) Dalton's Law (b) Lavosier (c) Faraday's law (d) None of these.

28. What is range of weak acids in PH scale:

(a) 3 to 4 (b) 3 to 5 (c) 1 to 3 (d) 2 to 4

29. The Formula of sugar is:

(a) $C_{12}H_{22}O_{12}$ (b) $C_{12}H_{24}O_{12}$ (c) $C_{11}H_{22}O_{11}$ (d) $C_{12}H_{22}O_{11}$

30. The formula of Chloroform:

(a) $CHCl_3$ (b) CH_3Cl (c) CH_2Cl_2 (d) All of these

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SECTION "B"

SHORT QUESTIONS

NOTE: Attempt any SIX questions each question carry 3 marks. /18

Q.No: 2 Define law of Multiple proportion with examples?

Q.No:3 Explain main point of Rutherford Atomic Model?

Q.No: 4 What is Brownian Movement ?

Q.No:5 Explain Covalent Bond with its types?

Q.No:6 Write definition of Metals, Non-Metals and Metalloid With examples?

Q.No:7 What would be the P_H of 0.001 M solution of HCl?

Q.No:8 Explain Faraday's First law of electrolysis?

Q.No:9 Calculate the formula mass of given:

(a) KNO_3 (b) $C_{12}H_{22}O_{11}$

(c) $CHCl_3$ (d) $MgCl_2$

Q.No:10 Balance the Equation by inspection method.

i. $NH_3 + O_2 \longrightarrow NO + H_2O$

ii. $CaCO_3 + HCl \longrightarrow CaCl_2 + H_2O + CO_2$

iii. $KNO_3 \longrightarrow KNO_2 + O_2$

iv. $NaHCO_3 \longrightarrow Na_2CO_3 + H_2O + CO_2$

SECTION "C"

LONG QUESTIONS

NOTE: Attempt any TWO questions each question carry 6 marks. /12

Q.No:11 (A) Define Chemical Reaction and its types with suitable examples?

(B) Write at least four strong acids, weak acids, strong base and weak base?

Q.No:12 (A) Differentiate between Covalent bond and Co-ordinate Covalent Bond?

(B) Calculate the morality of solution containing 16gm glucose per 300 mL solution.

Q.No:13 (A) Define the role of chemistry in our society?

(B) Calculate the amount of silver deposited when 10 ampere of current is passed for 50 minutes through a solution of $AgNO_3$. (Z of Ag = 0.00118 g/C)